

REPORT DOCUMENTATION PAGE

AFRL-SR-AR-TR-04-

Public reporting burden for this collection of information is estimated to average 1 hour per response, including the time for reviewing instructions, searching existing data sources, gathering the required data, completing and reviewing this collection of information. Send comments regarding this burden estimate or any other aspect of this collection of information, including suggestions for reducing this burden, to Washington Headquarters Services, Directorate for Information Operations and Reports (0704-0188), Paperwork Project, Washington, DC 20540-6001. Respondents should be aware that notwithstanding any other provision of law, no person shall be subject to a penalty for failing to provide information unless it is specifically required by law. **PLEASE DO NOT RETURN YOUR FORM TO THE ABOVE ADDRESS.**

0159

1. REPORT DATE (03-11-2004)		2. REPORT TYPE Final Technical Report		3. DATES COVERED (From - To) Dec 1, 2000-Nov.30 2003	
4. TITLE AND SUBTITLE Large amplitude vibrational influence on electron transfer reactions				5a. CONTRACT NUMBER F49620-01-1-0193	
				5b. GRANT NUMBER F49620-01-1-0193	
				5c. PROGRAM ELEMENT NUMBER	
6. AUTHOR(S) A.M. Wodtke				5d. PROJECT NUMBER	
				5e. TASK NUMBER	
				5f. WORK UNIT NUMBER	
7. PERFORMING ORGANIZATION NAME(S) AND ADDRESS(ES) University of California Santa Barbara Department of Chemistry and Biochemistry, University of California Santa Barbara, Santa Barbara, CA 93106				8. PERFORMING ORGANIZATION REPORT NUMBER	
9. SPONSORING / MONITORING AGENCY NAME(S) AND ADDRESS(ES) AirForce Office of Scientific Research				10. SPONSOR/MONITOR'S ACRONYM(S)	
				11. SPONSOR/MONITOR'S REPORT NUMBER(S)	
12. DISTRIBUTION / AVAILABILITY STATEMENT Approve For Public Release; Distribution Unlimited.					
13. SUPPLEMENTARY NOTES					
14. ABSTRACT Our research into the chemical properties of highly vibrationally excited molecules has revealed a strong connection between large amplitude vibrational motion and the ability of a molecule to accept or donate electrons. In this proposal, we explain the work leading up to this insight and suggest ways to further test our understanding of this phenomenon. We suggest two experiments to extrapolate these results to completely new kinds of chemical encounters: the vibrational enhancement of electron capture and the direct observation of vibrational promotion of an electron transfer reaction. This work may have significant implications for electron scavenging around re-entry vehicles and accompanying communications blackouts.					
15. SUBJECT TERMS Highly Vibrationally Excited Molecules, Plasma, Infrared Signatures, Molecular Dynamics					
16. SECURITY CLASSIFICATION OF:			17. LIMITATION OF ABSTRACT	18. NUMBER OF PAGES	19a. NAME OF RESPONSIBLE PERSON A.M. Wodtke
a. REPORT	b. ABSTRACT	c. THIS PAGE			19b. TELEPHONE NUMBER (include area code) 805 893 8085

20040319 114

Final Technical Report for AFOSR Grant No. F49620-01-1-0193

Alec M. Wodtke

Department of Chemistry University of California, Santa Barbara, CA 93106

Overview

This document reports on work performed under the grant: *F49620-01-1-0193, Air Force Office of Scientific Research, \$455,000, Dec. 1, 2000-Nov. 30, 2003*. At the outset of the last funding period, our main interest concerned the scattering of highly vibrationally excited molecules from solid surfaces and the microscopic dynamics of such processes. One of our principal results concerns the ability of vibrational motion to promote electron transfer from the metal to the incoming molecule when highly vibrationally excited molecules collide at metal surfaces. An update on our study of vibrational promotion of electron transfer is presented below.

As part of this work we have been motivated to work on new ways of producing intense molecular beams of highly vibrationally excited molecules. To this end we have taken up efforts to transversely refocus molecular beams of highly vibrationally excited molecules using an electric hexapole field. The AFOSR has supported design and construction of a new apparatus in our laboratory that combines stimulated emission pumping of NO, and thereby state-selective optical excitation of highly vibrationally excited molecules, with hexapole focusing. This instrument has now been completed and

recently, we have succeeded in transversely refocusing NO($v=18$). Our preliminary results are described below.

One of the ongoing objectives of our work has been to find other conditions where vibrational promotion of electron transfer might be important. Specifically, we have begun working toward finding evidence of similar electron transfer dynamics in bimolecular gas-phase collisions. In the course of commissioning a velocity map imaging apparatus for these studies, we have performed experiments on the photochemistry of ClN_3 , which is the commonly used precursor to $\text{NCl}(a^1\Delta)$, an important energy reservoir compound for the chemical iodine laser. The results were very exciting. **Velocity map imaging results on ClN_3 photochemistry provide strong evidence of photochemical production of cyclic- N_3 .** These experiments represent the best available evidence that ring formation is possible in an all-Nitrogen molecule.

The remainder of this report is organized as follows. First, we present a list of publications and scientific presentations of the work supported by this grant. A brief review of the important results from our surface scattering and hexapole focusing experiments is presented. Finally, a rather detailed presentation of the new photochemical results on ClN_3 is presented the describes the most exciting work.

Publications Resulting From This Work

1. *Observation of vibrational excitation and de-excitation for NO($v=2$) scattering from Au(111): Evidence for electron-hole pair mediated energy transfer*, Y. Huang, A.M. Wodtke C.T. Rettner, and D.J. Auerbach, Phys. Rev. Lett. **84** 2985-8 (2000)
2. *Vibrational Promotion of electron transfer*, Y. Huang, S.J. Gulding, C.T. Rettner,

D.J. Auerbach A.M. Wodtke, Science **290**, 111-114 (2000)

3. *The Dynamics of "Stretched Molecules": Experimental Studies of Highly Vibrationally Excited Molecules With Stimulated Emission Pumping*, Michelle Silva, Rienk Jongma, Robert W. Field Alec M. Wodtke, Annual Reviews of Physical Chemistry **52**, 811-852 (2001)
4. *Chemistry with stretched molecules*, Alec Wodtke, Phys. Chem. Earth (C) **26**(7), 467-471 (2001)
5. *Non-Arrhenius Surface Temperature Dependence in Vibrational Excitation of NO on Au(111): Evidence for the Importance of Surface Electronic States*, Alec M. Wodtke, Huang Yuhui, Daniel J. Auerbach, Chemical Physics Letters **364**(3-4) 231-236 (2002).
6. *Interaction of NO($v=12$) with LiF(001): Anomalously large vibrational relaxation rates*, Alec M. Wodtke and Yuhui Huang, Daniel J. Auerbach, Journal of Chemical Physics (accepted)
7. *Transport and focusing of highly vibrationally excited NO molecules*, Daniel Matsiev, Jun Chen, M. Murphy, A.M. Wodtke Journal of Chemical Physics (submitted)
8. *Photodissociation Dynamics of ClN_3 at 203 nm: The NCl ($a^1\Delta / X^3\Sigma$) Product Branching Ratio*, N. Hansen and A. M. Wodtke, A. V. Komissarov and M. C.

Heaven, Chemical Physics Letters **368** 568–573 (2003).

9. *Ion Dissociation Dynamics of the Chlorine Azide Cation (ClN_3^+) Investigated by Velocity Map Imaging*, N. Hansen, A. V. Komissarov, K. Morokuma, M. C. Heaven and A. M. Wodtke, Journal of Chemical Physics (accepted)
10. *The Simplest All-Nitrogen Ring*, Nils Hansen, A.M. Wodtke, Science (submitted)
11. *The $\text{Cl} + \text{N}_3$ channel in the UV photochemistry of ClN_3* , N. Hansen and A. M. Wodtke, Journal of Physical Chemistry (Charles Parmenter Festschrift) (submitted).
12. *Rotational Analysis of the origin and the inversion bands in the $S_1 \leftarrow S_0$ spectrum of acetaldehyde*, E. Jalviste, G. Berden, M. Drabbels, A. M. Wodtke, J. Chem. Phys. **114**(19) 8316-27 (2001)
13. *A new Approach to ^{13}C isotope separation*, Nils Hansen and Alec M. Wodtke, Chemical Physics Letters **356**:3-4, 340-346 (2002)
14. *Invited Chapter for Advanced Series in Physical Chemistry, World Scientific, Chuik Ng Series Editor, Modern Trends in Chemical Reaction Dynamics Ed.'s Xueming Yang and Kopin Liu*, "Interactions of vibrationally excited molecules at surfaces: A probe for electronically non-adiabatic effects" Alec M. Wodtke (submitted).

Scientific Presentations of the Work Supported by This Grant

1. Invited Lecture, Department of Chemistry Colloquium Series, Emory University, Feb. 10, 2000.
2. Invited Lecture, Department of Chemistry Chemical Physics Seminar Series, California Institute of Technology, April 4, 2000.
3. Invited Lecture, FOM for Plasma Physics Rijnhuizen, The Netherlands, April 26, 2000
4. Invited Lecture, European Geophysical Society Meeting, Nice France, April 28, 2000.
5. Invited Lecture, Department of Chemistry EPFL, Lausanne Switzerland, May 1 2000.
6. Invited Lecture, Department of Physics University of Bielefeld, Bielefeld, Germany May 3, 2000.
7. Invited Lecture, AFOSR Contractors Meeting, Waltham, MA, May 22, 2000.
8. Invited Lecture, Institute of Atomic and Molecular Science (IAMS), Taipei Taiwan, June 12, 2000
9. Invited Lecture, Synchrotron Radiation Research Center, Hsinchu, Taiwan, June 13, 2000

10. Invited Lecture, CECAM Workshop on Molecular Aspects of Gas Dynamics, Lyon France, June 28
11. Invited Lecture, University of Kiel, Physical Chemistry Colloquium, Kiel Germany, June 29 2000.
12. Invited Lecture, Pacifichem 2000, Symposium on "Photon and Electron Induced Processes on Surfaces", December 14 - 19, 2000, Honolulu, Hawaii
13. Poster Presentation: Gordon Research Conference on Molecular Energy Transfer, 1/14/2001-1/19/2001, Ventura California
14. Invited Speaker, ACS meeting Orlando Florida, April 7 2002.
15. Invited Speaker, AFOSR Contractors Meeting, May 19-22, 2002, Waltham Massachusetts.
16. Invited Speaker, Atomic and Molecular Interactions Gordon Conference, July 7 2002.
17. Invited Speaker, "Physical Chemistry Seminar", Boston University, Boston Massachusetts, Sept. 16 2002
18. Invited Speaker "Physical Chemistry Seminar", Massachusetts Institute of Technology, Boston Massachusetts, Sept. 17 2002

19. Invited Speaker "Physical Chemistry Seminar", University of Sherbrooke, Quebec
Canada, Sept. 18 2002
20. Invited Speaker "Physical Chemistry Seminar", Emory University, Atlanta Georgia,
Sept. 20, 2002
21. Invited Speaker "Conference on Stereo-Dynamics of Chemical Reactions" Schoorl,
the Netherlands Dec. 1 – 6 2002
22. Invited Colloquium Speaker, University of Leiden, Leiden, the Netherlands Dec. 10,
2002
23. Invited Speaker, "Meeting of the European Network on Cold Molecule Science",
Rijnhuizen, Holland Dec. 11-12 2002.
24. Plenary Lecture, "International Symposium on Molecular Beams", Lisbon, Portugal,
June 8-13 200

Hexapole Focusing and Surface Scattering of Highly Vibrationally Excited

NO

One of the major motivations for our work concerns the properties of highly vibrationally excited Nitric Oxide. This molecule produces an important source of non-thermal infrared radiation in the upper atmosphere that is controlled by molecular interactions in both gas-phase bimolecular and molecule-surface collisions. Because this

molecule is found far out of local thermodynamic equilibrium, it is necessary to understand its vibrational state-specific dynamics. We have in the past worked hard to understand bimolecular gas-phase collisions involving highly vibrationally excited NO. More recently we have focused on collisions of highly vibrationally excited NO with solid surfaces.

*Electron Mediated Vibrational Energy Transfer at Metal Surfaces*¹

One of the most significant results of our work is the observation of efficient multi-quantum vibrational relaxation of NO($v=15$) in specular collisions with a gold surface. See Fig. 1. This has been explained by vibrational promotion of electron transfer. In this picture, large amplitude vibrational motion strongly influences electron transfer dynamics as the stretched NO bond may easily accept electrons, while the compressed NO bond repels them (Fig. 2). This enables the stretched NO molecule to accept an electron near the Fermi-level of the metal and upon bond compression deliver it back to the metal at a far higher energy converting large amounts of NO-vibration to metal-electron excitation. These insights have lead us to carry out experiments that probe more deeply models of electron mediated vibrational energy transfer. In two studies we have carried out experiments that push the limits of the present model of vibrational excitation of molecules at metal surfaces.

In the first, measurements of vibrational excitation and de-excitation of NO($v=2$) scattering from a Au(111) surface showed that the probability of both processes increases strongly with the kinetic energy of the incident NO. This qualitative result is consistent with a vibrational energy transfer mechanism involving electron-hole pairs but not with

¹ See the publication list: #1–#5.

competing adiabatic models. The form of the energy dependence and other features of the measurements, are however not quantitatively in accord with existing calculations.

In the second study, we reported the surface temperature dependence for $\Delta v=+1$ vibrational excitation of $\text{NO}(v=2)$ due to collisions with $\text{Au}(111)$ at two incidence energies. This was compared to previous measurements on $\text{Ag}(111)$ and $\text{Cu}(110)$. The standard model for electron-hole pair mediated vibrational excitation predicts that all three metals will exhibit a nearly Arrhenius temperature dependence with the same effective activation energy. While scattering on $\text{Ag}(111)$ and $\text{Cu}(110)$ obey this prediction, our new results show that the temperature dependence is non-Arrhenius on $\text{Au}(111)$.

We suggested that the large spin-orbit interaction present for $\text{Au}(111)$ surface electronic states may be important for vibrational excitation. These experiments indicate that the simple bulk electronic structure models of the dynamics of vibrational energy transfer at metals may be too simple.

*Vibrational Energy Transfer at Insulators*²

To compare the behavior on metals to that on insulators, we have made measurements of the vibrational survival probability of $\text{NO}(v=12)$ suffering collisions with a $\text{LiF}(001)$ crystalline surface. Survival probabilities are near unity in all cases, showing the key observation associated with vibrational promotion of electron transfer is absent on an insulator. At the lowest reported incidence energies (5.3 kJ/mol) and surface temperatures (290K), vibrational relaxation is enhanced, scattering angular distributions approach $\text{Cos}(\theta)$ and rotational temperatures of the scattered molecules approach the

surface temperature. At higher incidence energies, rotational temperatures are higher than the surface temperature and angular distributions narrow approaching $\cos^6\theta$. These results suggest that trapping/desorption is occurring at least at the lowest incidence energies and surface temperatures. Substantial vibrational relaxation appears to occur when residence time are as small as a few ps. This is many orders of magnitude faster than has been reported for diatomic molecules relaxing from $v=1$ on salt crystals. These experiments suggest that theoretical work is still needed to help us understand the vibrational relaxation of highly vibrationally excited molecules at insulator surfaces.

*Hexapole Focusing of NO($v=18$)*³

We have also obtained AFOSR support to develop a new means of producing highly vibrationally excited molecules in a molecular beam. This approach takes advantage of the state-specific nature of hexapole focusing and exploits this in combination with optical preparation of highly vibrationally excited NO. We have recently carried out experiments where hexapole focusing is combined with stimulated emission pumping in a molecular beam, providing control over the molecule's rovibronic quantum numbers, its laboratory frame velocity and its transverse divergence. Hexapole focusing profiles can be quantitatively reproduced by classical trajectory simulations. These experiments provide new ways of manipulating beams of vibrationally excited molecules including: 1) transverse refocusing and concomitant improved efficiency for transport of the vibrationally excited molecules, 2) relative enrichment of the concentration of the vibrationally excited molecules with respect to the

² See the publication list: #6.

³ See the publication list: #7.

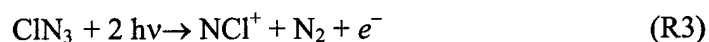
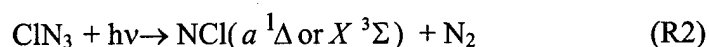
unexcited portion of the beam and, 3) orientation of vibrationally excited molecules. See Fig. 3.

The remainder of this report concerns velocity map imaging investigations into the photochemistry of ClN_3 . This work has led us to conclude that ClN_3 is a photochemical precursor to cyclic- N_3 , which will be the focus of the next funding period. Thus the following sections serve to introduce the topic of cyclic- N_3 production as well as present our prior results.

Production of Cyclic- N_3 : Velocity Map Imaging of ClN_3 Photochemistry⁴

Overview of ClN_3 Photochemistry

Using a new velocity map imaging apparatus developed under the support of the AFOSR, we (in collaboration with Prof. Michael Heaven⁵ of Emory University) have recently carried out the first collision-free experiments on the ultraviolet photochemistry of chlorine-azide (ClN_3). We have studied three photochemical channels:



Recent interest in the photochemistry of chlorine azide (ClN_3) derives from

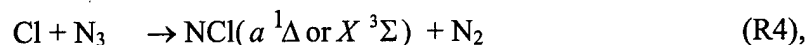
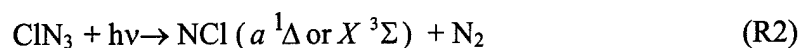
demonstrations that the primary photoproduct $\text{NCl}(a^1\Delta)$ can be used as an effective energy carrier in chemical iodine lasers, a topic of significant relevance to the AFOSR

⁴ See the publication list: #8 – #11.

⁵ Professor Heaven and Dr. Anatoly Komissarov spent about two weeks teaching us the use of the ClN_3 generator used to form molecular beams of this molecule and after returning to Atlanta have been very

mission (1-3). This stimulated several past studies of the near UV photolysis of ClN_3 , in particular by Dehnicke *et al.* (4), Coombe and co-workers (5-9), Henshaw *et al.* (10), and Komissarov *et al.* (11, 12).

Despite substantial effort, basic features of this chemistry still had not been established prior to our work. For example, the thermochemistry of the two most important reactions used to form $\text{NCl}(a^1\Delta)$:



remained unclear, as did the singlet/triplet branching ratio for both reactions. In our experiments, we were able to obtain the maximum release of translational energy from the velocity map image of recoiling REMPI-detected N_2 in reaction R2. This allowed us to obtain an accurate experimental value for this reaction's thermochemistry⁶ and led to a measure of the singlet/triplet branching ratio for reaction R2 (13) that supported prior experiments done under collisional conditions (12).

Velocity map imaging experiments on all of the products of reaction R3 led to accurate thermochemistry for this reaction. Through knowledge of the ionization potentials of NCl (14) and ClN_3 (15) an independent determination of the thermochemistry of reaction R2 was also possible. The two independent experiments were in quantitative agreement with one another on the 0.1 eV level of accuracy. The

helpful in performing some theoretical calculations and performing some complementary experiments not described in this report.

⁶ The mechanics of this kind of analysis is presented in more detail below when we come to a consideration of the evidence supporting the photochemical production of cyclic- N_3 .

latter work also provided a detailed look at the dynamics of the decomposition of the unstable ClN_3^+ ion (16).

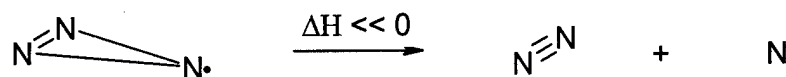
By deriving the maximum release of translational energy for Cl-atoms recoiling from reaction R1, we were likewise able to experimentally determine the bond energy, $D_0(\text{Cl}-\text{N}_3)$, (17, 18). This thermochemistry could be combined with the thermochemistry of reaction R2 to obtain accurate thermochemistry of the second important $\text{NCl}(a^1\Delta)$ forming reaction R4.

The results of our new experiments, together with past work now provides accurate (within 0.1 eV) thermochemistry for all molecules containing one Cl-atom and three Nitrogen-atoms, providing a better foundation for understanding the chemistry of $\text{NCl}(a^1\Delta)$. The derived thermochemistry is summarized in Fig. 4.

Photochemical Production of Cyclic- N_3

In the course of these studies, we discovered what is certainly the most significant result of the last funding period, both from a fundamental point of view and from the point of view of AFOSR mission. **Our experiments provide strong evidence that ClN_3 is a photochemical precursor of the simplest all-Nitrogen cyclic molecule, cyclic- N_3 .**

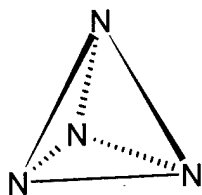
For Nitrogen unlike most elements, the energy of an N-N single bond is substantially less than one-third, and the energy of an N=N double bond is substantially less than two-thirds of an $\text{N}\equiv\text{N}$ triple bond. Consequently, all-Nitrogen species that form rings at the expense of multiple bonds are subject to a strong chemical driving force toward dissociation to N_2 .



Not surprisingly, the ability of Nitrogen to form cyclic structures has been difficult to verify. Several experiments have led to the postulate of a cyclic poly-nitrogen species; but these have either remained unconfirmed or later been disproved (19-22). Recently, the stabilization energy of an ionic crystal was exploited, allowing production of N_5^- in an ionic salt (23). Still the possible existence of isolated cyclic all-Nitrogen allotropes has remained largely a topic for theory (22, 24-41) and the question of how such molecules might be produced in the laboratory has remained unanswered (42).

Partly due to the fact that this class of molecules is so unstable, they have been vigorously sought as materials capable of storing large amounts of energy in minimal volumes, so-called "high energy and density materials" (43). Their natural tendency to decompose to N_2 also makes them high on the list of environmentally friendly explosives and propellants.

The most studied example, tetraazahedrane (19, 22, 24, 25, 28-35) better known as tetrahedral Nitrogen, **molecule (1)**, is illustrative of the chemical properties of this family of molecules.

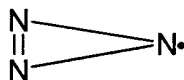


Molecule (1)

Formally in compliance with the octet-rule, it is a closed shell-species of singlet spin-multiplicity (24) that may release more than 760 kJ/mole by decomposition to 2 N_2 (22,

28). Theory shows that the barrier to decomposition to ground state N_2 is greater than 250 kJ/mole, strongly suggesting that it is an isolable molecule (28, 30). However, decomposition via "spin-forbidden" curve-crossings forming electronically excited products, which is predicted to occur over a smaller barrier (31), may lower the stability of **molecule (1)**. **Molecule (1)**, like all other gas-phase cyclic poly-N molecules, has yet to be observed (19, 22).

The simplest all-Nitrogen allotrope from which a ring can be formed is N_3 (25-27). Recent high-level calculations (25) predict a C_{2v} molecule of B_1 symmetry and doublet spin-multiplicity with two long (1.455 Å) and one short (1.218 Å) N/N bonds (44), **molecule (2)**.



Molecule (2)

Isomerization to the well-known linear isomer (the azide radical) is predicted to be exoergic by 125 kJ/mole, spin-allowed but limited by a 140 kJ/mole barrier (25).

Dissociation to ground state $N(^4S_{3/2})$ and $N_2(\tilde{X}^1\Sigma_g^+)$ is exoergic by ~125 kJ/mole, but spin-forbidden (25). The lowest energy spin-allowed dissociation forms $N(^2D)$ and $N_2(\tilde{X}^1\Sigma_g^+)$ and is endoergic by ~96 kJ/mole (25). Theoretical calculations also suggest that cyclic- N_3 may undergo barrierless exoergic recombination with $N(^2D)$ to form tetrahedral N_4 (25). **This makes cyclic- N_3 one of the only suggested precursors to tetrahedral N_4 (39).**

In the course of these studies, we have measured accurate velocity maps of Cl-atoms formed in reaction R1. Like photoelectron spectroscopy where the translational energy distribution of the electron reveals quantum-level structure of the simultaneously formed ion, velocity-map images of quantum-state-selected Cl-atoms recoiling from photolyzed ClN_3 reveal a medium resolution energy spectrum of N_3 . These experiments indicate the presence of a low-lying excited state of N_3 at 1.35 ± 0.1 eV, which compares well with recent theoretical predictions of the energy of cyclic- N_3 , **molecule (2)** (25).

Fig. 5a shows the observed velocity map image of ground-state Cl-atoms produced in reaction (R1). Here, the direction of the ClN_3 molecular beam is out of the plane of the paper pointing directly at the reader. The polarization direction of the laser light beam is also indicated (\hat{e}). The position at which a Cl^+ ion is detected in the image reflects the velocity (both direction and speed) of the recoiling Cl-atom from which it was formed.

One will immediately notice that there is a limiting speed (indicated by the dotted circle) beyond which no Cl-atoms are detected. This limit is found when the maximum possible fraction of the absorbed photon energy beyond that required to break the N-Cl bond in ClN_3 is channeled into translational energy of the photoproducts, leaving N_3 with a minimum of internal energy. This limiting velocity provides a means of obtaining an upper limit to the N-Cl bond energy in ClN_3 , $D_0(\text{Cl}-\text{N}_3)$. It can be shown that this velocity is given by Eq. (1).

$$v_{\text{Cl}}^{\text{MAX}} \leq \sqrt{[h\nu - D_0(\text{Cl} - \text{N}_3)] \frac{2m_{\text{N}_3}}{m_{\text{Cl}}(m_{\text{N}_3} + m_{\text{Cl}})}} \quad \text{Eq. (1)}$$

where, $h\nu$ is the energy of the photon absorbed by ClN_3

and $D_0(\text{Cl-N}_3)$ is the N-Cl bond energy.

The vast majority of the photodissociation events, where some internal energy appears in the N_3 photofragment, appear well within the dotted circle. Also visible in Fig. 5a is a second limiting speed (indicated by the dashed circle) beyond which *many* Cl-atoms are never formed. Since only a single quantum state of the Cl-atoms are detected with REMPI, this indicates that the internal energy distribution of the N_3 photofragment is markedly bimodal.

To analyze this quantitatively, we derived the velocity and angular distributions of the Cl-atoms from an inverse-Abel transform of the image of Fig. 5a (45). The corresponding translational energy distribution of the recoiling Cl-atoms is shown in Fig. 5b. One sees now clearly the translational energy cut-off at a Cl-atom translational energy of 1.65 eV. We may use Eq. (1) to obtain Eq. (2) below:

$$D_0(\text{Cl} - \text{N}_3) \leq h\nu - \frac{1}{2} m_{\text{Cl}} (v_{\text{Cl}}^{\text{MAX}})^2 \frac{m_{\text{N}_3} + m_{\text{Cl}}}{m_{\text{N}_3}} \quad \text{Eq. (2)},$$

and derive the dissociation energy of ClN_3 , $D_0(\text{Cl-N}_3) = 2.24 \pm 0.1$ eV, which is in agreement with other thermodynamic quantities (46-49), in particular the ClN_3 heat of formation. This shows that the fastest detected Cl-atoms correspond to formation of the linear azide radical with little or no internal excitation.

In a way reminiscent of photoelectron spectroscopy, it is useful to realize that the Cl-atom translational energy distribution shown in Fig. 5b is likewise a medium resolution energy spectrum of the N_3 fragment formed in the photodissociation or ClN_3 . By the principle of energy conservation, low translational-energy Cl-atoms directly correspond to high internal energy N_3 radicals. To make this clear, the upper x-axis of Fig. 5b shows how the measurement reflects the N_3 internal energy distribution. The most

remarkable observation of this work is the bimodal nature of this distribution, with the origins of the two components appearing at 1.65 and 0.9 eV in Cl-atom translational energy space, respectively. Alternatively, one may state that the origins of the two components appear at 0.0 and 1.35 eV internal energy of the N_3 molecule, the first of which is assigned to the linear azide radical.

An energy gap of 1.35 eV would normally be attributed to electronic excitations in a molecule; however, there is no experimental evidence of an excited electronic state of the well-studied linear azide N_3 lying 1.35 eV above the ground state⁷. Extensive spectroscopic studies have revealed the presence of an excited state of $^2\Sigma$ -symmetry at 4.5 eV (50) and theoretical work shows that the lowest linear quartet is also at about 4.4 eV (27). Recently, the electronic structure of N_3 has been investigated in more depth using quantum chemical methods (25). This study has shown that the experimentally observed $^2\Sigma$ -state is the lowest lying excited doublet electronic state with a *linear* structure and the calculated energy is in good agreement with experiment (50). These same theorists also predict that the lowest excited electronic state of N_3 is a ring structure, belonging to the C_{2v} -point group and possessing 2B_1 symmetry, **molecule (2)**. At the highest levels of theory, these authors found the excitation energy to this state (which is the cyclization energy of N_3) to be 1.35 ± 0.05 eV⁸ (25). This is shown as the red bar in Fig. 5b. The good agreement with the derived origin of the high energy component of the N_3 excitation distribution is compelling evidence that cyclic- N_3 is formed in the ultraviolet photochemistry of ClN_3 .

⁷ Preliminary theoretical calculations from Morokuma's group give no evidence of a bimodal vibrational energy distribution in the N_3 fragment, a result that would be in any event quite surprising.

⁸ This includes a zero-point energy correction.

References

1. A. J. Ray, R. D. Coombe, *J. Phys. Chem.* **99**, 7849-7852 (1995).
2. T. L. Henshaw, G. C. Manke, T. J. Madden, M. R. Berman, G. D. Hager, *Chem. Phys. Lett.* **325**, 537-544 (2000).
3. G. C. Manke, G. D. Hager, *J. Mod. Opt.* **49**, 465-474 (2002).
4. K. Dehnicke, P. Ruschke, *Z.Naturforsch.(B)* **33**, 750-752 (1978).
5. A. J. Ray, R. D. Coombe, *J. Phys. Chem.* **98**, 8940-8945 (1994).
6. R. D. Coombe, D. Patel, A. T. Pritt, F. J. Wodarczyk, *J. Chem. Phys.* **75**, 2177-2190 (1981).
7. R. D. Coombe, S. J. David, T. L. Henshaw, D. J. May, *Chem. Phys. Lett.* **120**, 433-436 (1985).
8. R. D. Coombe, M. H. Vanbentham, *J. Chem. Phys.* **81**, 2984-2990 (1984).
9. R. H. Jensen, A. Mann, R. D. Coombe, *J. Phys. Chem. A* **104**, 6573-6579 (2000).
10. T. L. Henshaw, S. D. Herrera, G. W. Haggquist, L. A. V. Schlie, *J. Phys. Chem. A* **101**, 4048-4056 (1997).
11. A. V. Komissarov, G. C. Manke II, S. J. Davis, M. C. Heaven, *Proc. SPIE - Int. Soc. Opt. Eng.* **3931**, 138-148 (2000).
12. A. V. Komissarov, G. C. Manke, S. J. Davis, M. C. Heaven, *J. Phys. Chem. A* **106**, 8427-8434 (2002).
13. N. Hansen, A. M. Wodtke, A. V. Komissarov, M. C. Heaven, *Chem. Phys. Lett.* **368**, 568-573 (2003).
14. V. Butcher, J. M. Dyke, A. E. Lewis, A. Morris, A. Ridha, *J. Chem. Soc. - Faraday Trans. II* **84**, 299-310 (1988).
15. D. C. Frost, C. B. Macdonald, C. A. McDowell, N. P. C. Westwood, *Chem. Phys.* **47**, 111-124 (1980).
16. N. Hansen, A. M. Wodtke, A. V. Komissarov, K. Morokuma, M. C. Heaven, *J. Chem. Phys. (submitted)* (2002).
17. N. Hansen, A. M. Wodtke, Z. Peng, K. Morokuma, *Science (submitted)* (2003).
18. N. Hansen, A. M. Wodtke, *J. Phys. Chem. A (in progress)* (2003).
19. J. P. Zheng, J. Waluk, J. Spanget-Larsen, D. M. Blake, J. G. Radziszewski, *Chem. Phys. Lett.* **328**, 227-233 (2000).
20. J. P. Calfa, K. G. Phelan, F. T. Bonner, *Inorg. Chem.* **21**, 521-524 (1982).
21. R. J. Gowland, K. R. Howes, G. Stedman, *J. Chem. Soc.-Dalton Trans.*, 797-799 (1992).
22. T. J. Lee, J. M. L. Martin, *Chem. Phys. Lett.* **357**, 319-325 (2002).
23. A. Vij, J. G. Pavlovich, W. W. Wilson, V. Vij, K. O. Christe, *Angew. Chem.-Int. Edit.* **41**, 3051-3054 (2002).
24. M. Bittererova, T. Brinck, H. Ostmark, *J. Phys. Chem. A* **104**, 11999-12005 (2000).
25. M. Bittererova, H. Ostmark, T. Brinck, *J. Chem. Phys.* **116**, 9740-9748 (2002).
26. R. Klein, S. Biskupic, *Chem. Pap.-Chem. Zvesti* **47**, 143-148 (1993).
27. J. Wasilewski, *J. Chem. Phys.* **105**, 10969-10982 (1996).
28. K. M. Dunn, K. Morokuma, *J. Chem. Phys.* **102**, 4904-4908 (1995).
29. T. J. Lee, C. E. Dateo, *Chem. Phys. Lett.* **345**, 295-302 (2001).
30. T. J. Lee, J. E. Rice, *J. Chem. Phys.* **94**, 1215-1221 (1991).

31. D. R. Yarkony, *J. Am. Chem. Soc.* **114**, 5406-5411 (1992).
32. H. Ostmark, O. Launila, S. Wallin, R. Tryman, *J. Raman Spectrosc.* **32**, 195-199 (2001).
33. M. N. Glukhovtsev, H. J. Jiao, P. V. Schleyer, *Inorg. Chem.* **35**, 7124-7133 (1996).
34. A. A. Korkin, A. Balkova, R. J. Bartlett, R. J. Boyd, P. V. Schleyer, *J. Phys. Chem.* **100**, 5702-5714 (1996).
35. W. J. Lauderdale, J. F. Stanton, R. J. Bartlett, *J. Phys. Chem.* **96**, 1173-1178 (1992).
36. M. Bittererova, H. Ostmark, T. Brinck, *Chem. Phys. Lett.* **347**, 220-228 (2001).
37. M. Bittererova, T. Brinck, H. Ostmark, *Chem. Phys. Lett.* **340**, 597-603 (2001).
38. M. Tobita, R. J. Bartlett, *J. Phys. Chem. A* **105**, 4107-4113 (2001).
39. M. L. Leininger, T. J. VanHuis, H. F. Schaefer, *J. Phys. Chem. A* **101**, 4460-4464 (1997).
40. A. Larson, M. Larsson, H. Ostmark, *J. Chem. Soc.-Faraday Trans.* **93**, 2963-2966 (1997).
41. M. T. Nguyen, T. K. Ha, *Chem. Berichte* **129**, 1157-1159 (1996).
42. F. Cacace, G. de Petris, A. Troiani, *Science* **295**, 480-481 (2002).
43. L. E. Fried, M. R. Manaa, P. F. Pagoria, R. L. Simpson, *Ann. Rev. Mater. Res.* **31**, 291-321 (2001).
44. Recent and still unpublished calculations from Morokuma's group suggest that there may be another nearly isoenergetic ring isomer of 2A2 symmetry
45. V. Dribinski, A. Ossadtchi, V. A. Mandelshtam, H. Reisler, *Rev. Sci. Instrum.* **73**, 2634-2642 (2002).
46. K. P. Huber, G. Herzberg, *Molecular Spectra and Molecular Structure* (Van Nostrand Reinhold Company Inc., New York, ed. 1, 1979), vol. IV. Constants of Diatomic Molecules.
47. R. E. Continetti, D. R. Cyr, R. B. Metz, D. M. Neumark, *Chem. Phys. Lett.* **182**, 406-411 (1991).
48. R. E. Continetti, D. R. Cyr, D. L. Osborn, D. J. Leahy, D. M. Neumark, *J. Chem. Phys.* **99**, 2616-2631 (1993).
49. C. Paillard, R. Moreau, Combouri.J, *Comptes Rendus Hebdomadaires Des Seances De L Academie Des Sciences Serie C* **264**, 1721-& (1967).
50. A. E. Douglas, W. J. Jones, *Can. J. Phys.* **43**, 2216-& (1965).

Figures and Captions

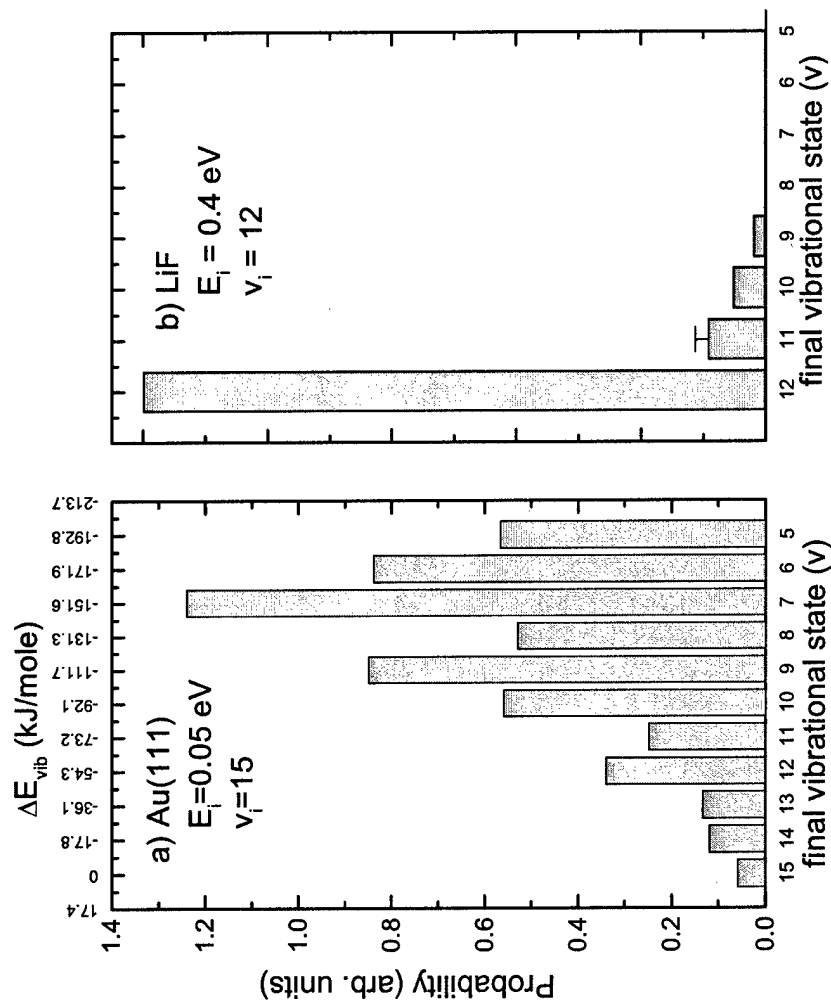


Figure 1. Measured vibrational distribution of NO resulting from scattering of a) NO($v=15$) from Au(111) at $E_{\text{incidence}} = 0.05 \text{ eV}$ and b) NO($v=12$) from LiF at $E_{\text{incidence}} = 0.4 \text{ eV}$. The remarkable multi-quantum vibrational relaxation seen on the metal has been explained in terms of vibrationally mediated electron transfer. From Y. H. Huang, C. T. Rettner, D. J. Auerbach, A. M. Wodtke, *Science* **290**, 111-114 (2000).

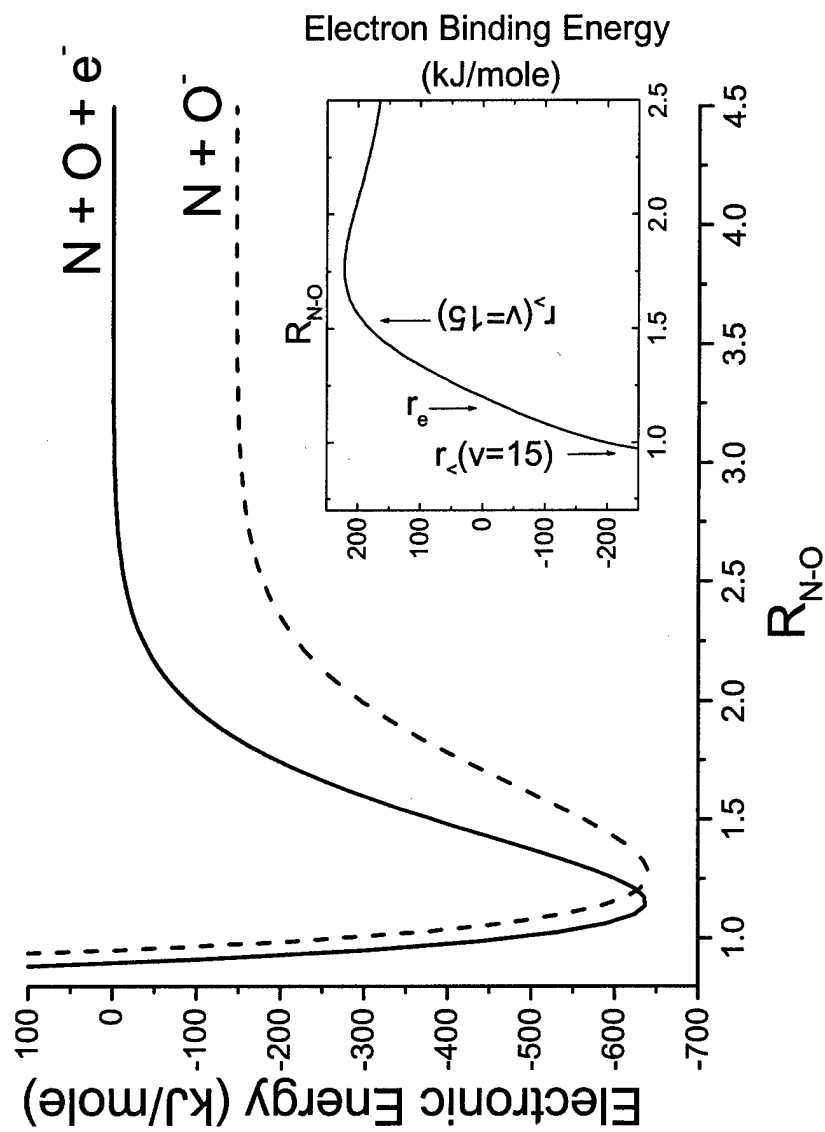


Figure 2: Ab initio calculations of NO and NO^- , showing the energetic constraints on electron transfers. The inset shows the difference between the two potential curves, the vertical electron binding energy. The positions of the inner and outer extrema of the NO $v=15$ vibrational wave function are shown with vertical arrows labeled (r_- and r_+). The position of the NO potential minimum is shown as r_e . This shows that the energetics of electron transfer depend strongly on the bond length. From Y. H. Huang, C. T. Rettner, D. J. Auerbach, A. M. Wodtke, *Science* **290**, 111-114 (2000).

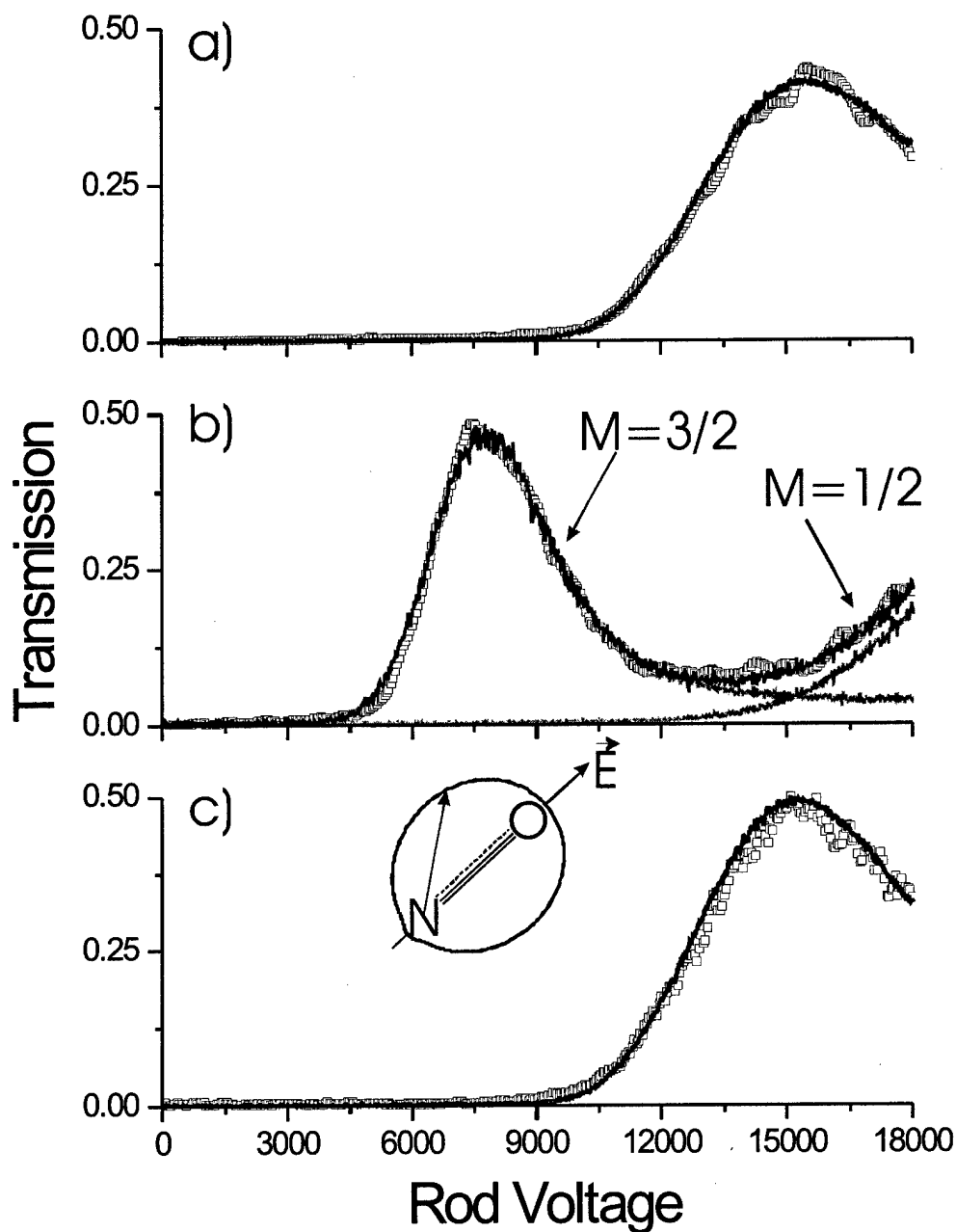


FIG. 3 Hexapole focusing curves for NO molecules pumped by stimulated emission pumping compared to ground state Molecules: (a) $|v=18, J=1/2, \Omega=1/2\rangle$, (b) $|v=18, J=3/2, \Omega=3/2\rangle$ and (c) $|v=0, J=1/2, \Omega=1/2\rangle$. Data is shown as open squares. The solid black lines are a classical trajectory simulation that quantitatively reproduces the data. The grey lines in (b) are the $M = 1.5$ and 0.5 magnetic sub-level components. This result demonstrates how optically prepared molecules delivered to an $\Omega=3/2$ level can be made to focus at a much lower voltage than the initially unexcited sample. Also shown as an inset is a polar plot of the oriented $|v=18, J=3/2, \Omega=3/2, M=3/2\rangle$ state of NO prepared at 8 kV Rod Voltage.

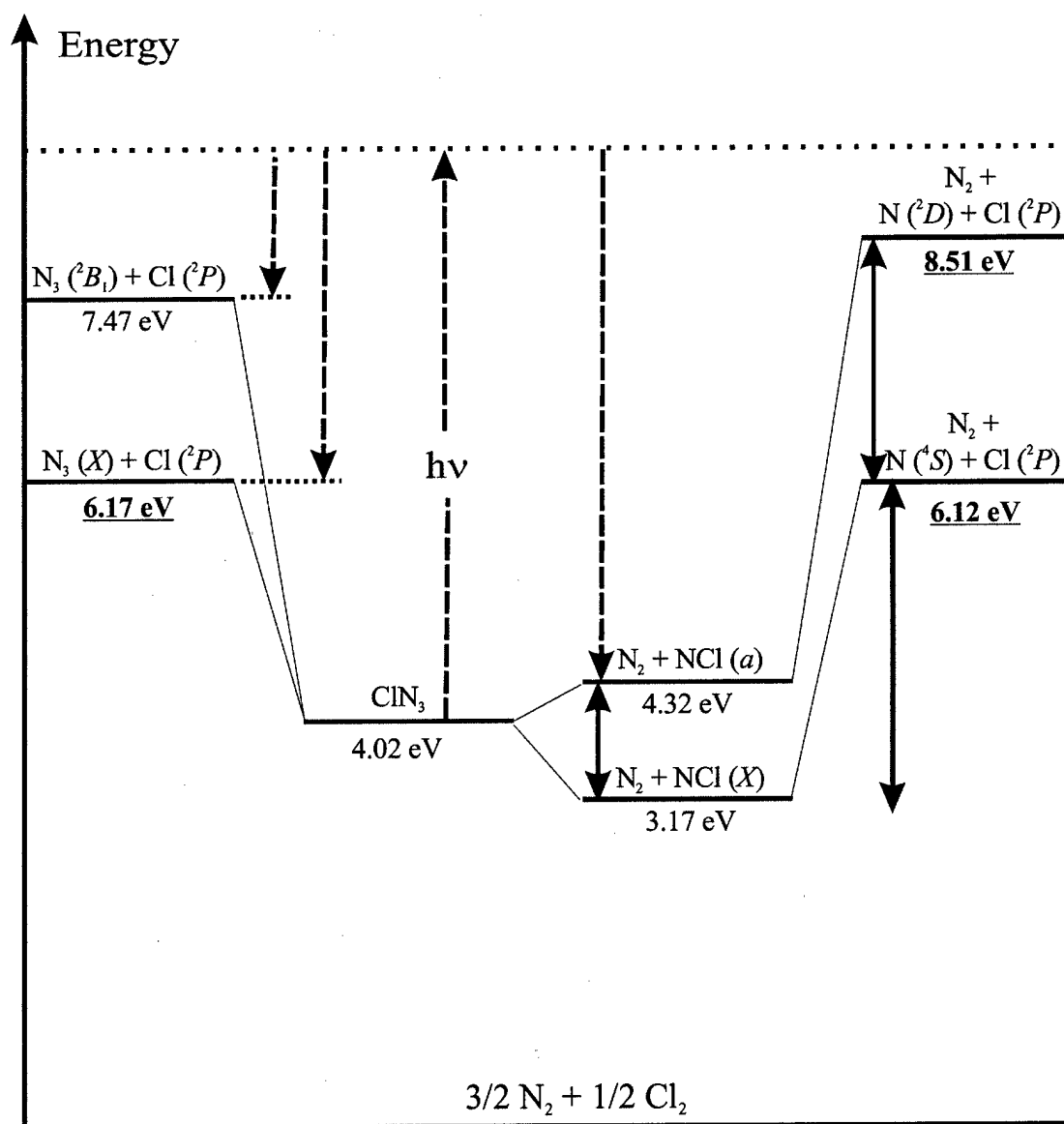


Figure 4: Zero Kelvin standard heats of formation (in eV) for molecules that may be constructed from a Cl-atom and three N-atoms. Constructed from recent velocity map imaging results together with past work. The bold underlined values indicate previously well-established heats of formation. The double headed arrows represent well-known thermodynamic transformations. The dashed, single-headed arrows represent thermodynamic transformations observed in our work. In addition, we have observed two-photon ionization (reaction R3) which adds redundant confirmation of these thermodynamics.

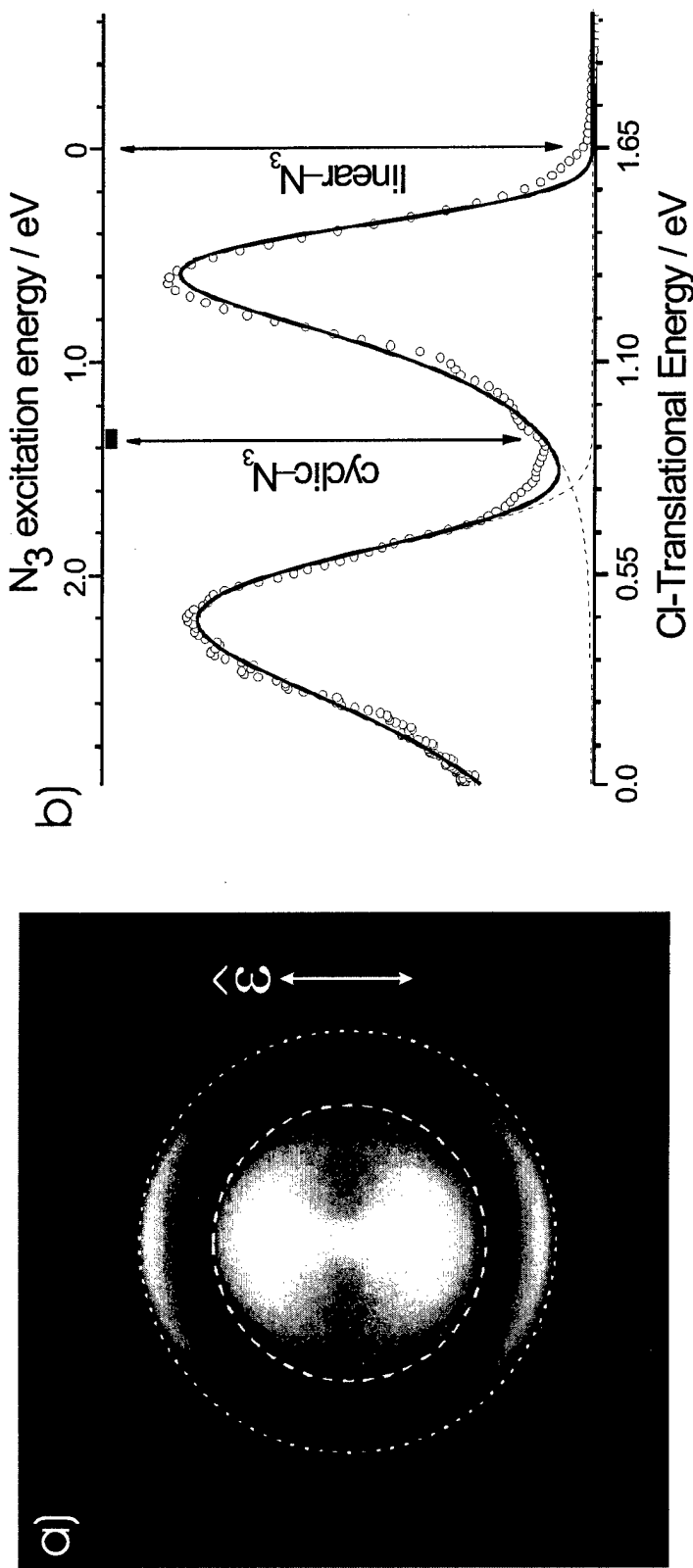


Figure 5: (a) Velocity map image of ground state $\text{Cl}(\text{P}_{3/2})$ recoiling from ClN_3 . The outer ring records Cl-atoms formed with linear azide radical with little internal energy. The inner ring records Cl-atoms formed with N_3 in an excited state. (b) The translational energy distribution for ground state $\text{Cl}(\text{P}_{3/2})$ derived from the velocity map image. The maximum release of translational energy is limited by the N-Cl bond strength in ClN_3 . This translational energy is indicated at 1.65 eV. At this Cl-atom translational energy, N_3 is formed with negligible internal excitation. Slower moving Cl-atoms reflect formation of more internally excited N_3 . The N_3 excitation energy can be read from upper x-axis. The analogy with photoelectron spectroscopy is obvious. The double-headed arrow marked "cyclic N_3 " shows where a new feature of the Cl-atom energy spectrum appears reflecting formation of an excited form of N_3 . The most recent theoretical predictions of the cyclization energy of N_3 are shown as the red bar. See text.